

$$0.93^{10} = e^{-0.93}$$

$$\text{MW NaOH } 32 \text{ g/mol} \quad \rho = 0.79 \text{ g/mL} \quad 10 \text{ mL } 7.9 \text{ g}$$

$$7.9 \text{ g} \frac{1 \text{ mol}}{32 \text{ g}} = 0.25 \text{ mol}$$

$$\frac{0.25 \text{ mol}}{12 \text{ mL}} \left(\frac{1000 \text{ mL}}{\text{L}} \right) = 20.8 \text{ M}$$

$$\frac{20.8}{3.5} = 5.94$$

$$\text{pH} = -\log [\text{H}^+]$$

$$\begin{array}{r} 0.083 \text{ } 0.083 \text{ } 1.1 \\ 0.417 \text{ } 0.4 \end{array}$$

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